

***SCHEME OF EXAMINATION
AND
SYLLABUS
(for Academic Session 2017-2018)***

**B.Sc. Chemistry
First & Second Semester Examination**

**This syllabus is only for the B.Sc. Course
running under Semester Scheme
at the University campus**

**Bachelor of Science (B.Sc.)
Chemistry**

Faculty of Science



UNIVERSITY OF KOTA

MBS Marg, KOTA (Rajasthan)-324 005

INDIA

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Bachelor of Science (B.Sc.) (Mathematics Group)

Subject Combination: Chemistry, Mathematics, Physics (CMP)

B.Sc. 1st Year : First and Second Semesters Semester Scheme of Examination

Year / Semester	Number, Code or ID and Nomenclature of Paper			Duration of Exam. (in Hrs.)	Teaching Hrs / Week & Credit Points			Distribution of Assessment Marks				Total Marks	
	Number of Paper	Code or ID of Paper	Nomenclature of Paper		Teaching Th.	Pr.	Credit Points	Continuous Assessment (20%)		Semester Assessment (80%)			
								Max. Marks	Min. Pass Marks	Max. Marks	Min. Pass Marks	Max. Marks	Min. Pass Marks
I Semester	Paper-1.1	HIND-111	General Hindi	2	2	-	2	-	-	50	20	50	20
	Paper-1.2	ECA-112	Elementary Computer Applications	2	2	-	2	-	-	50	20	50	20
	Paper-1.3	CHEM-111	Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-1.4	CHEM-112	Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-1.5	CHEM-113	Chemistry Practical	6	-	4	2	--	--	50	25	50	25
	Paper-1.6	MATH-111	Mathematics-I	3	3	-	3	15	06	60	24	75	30
	Paper-1.7	MATH-112	Mathematics-II	3	3	-	3	15	06	60	24	75	30
	Paper-1.8	MATH-113	Mathematics Practical	6	-	4	2	--	--	50	25	50	25
	Paper-1.9	PHY-111	Physics-I	3	3	-	3	15	06	60	24	75	30
	Paper-1.10	PHY-112	Physics-II	3	3	-	3	15	06	60	24	75	30
	Paper-1.11	PHY-113	Physics Practical	6	-	4	2	--	--	50	25	50	25
Total (I Semester)				36+4	30+4	24+4	90	36	510+100	219+40	600+100	195+40	
II Semester	Paper-2.1	ENG-121	General English	2	2	-	2	-	-	50	20	50	20
	Paper-2.2	ENV-122	Environmental Studies	2	2	-	2	-	-	50	20	50	20
	Paper-2.3	CHEM-121	Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-2.4	CHEM-122	Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-2.5	CHEM-123	Chemistry Practical	6	-	4	2	--	--	50	25	50	25
	Paper-2.6	MATH-121	Mathematics-I	3	3	-	3	15	06	60	24	75	30
	Paper-2.7	MATH-122	Mathematics-II	3	3	-	3	15	06	60	24	75	30
	Paper-2.8	MATH-123	Mathematics Practical	6	-	4	2	--	--	50	25	50	25
	Paper-2.9	PHY-121	Physics-I	3	3	-	3	15	06	60	24	75	30
	Paper-2.10	PHY-122	Physics-II	3	3	-	3	15	06	60	24	75	30
	Paper-2.11	PHY-123	Physics Practical	6	-	4	2	--	--	50	25	50	25
Total (II Semester)				36+4	30+4	24+4	90	36	510+100	219+40	600+100	195+40	

Bachelor of Science (B.Sc.) (Mathematics Group)

Subject Combination: Chemistry, Mathematics, Physics (CMP)

B.Sc. 2nd Year : Third and Fourth Semesters Semester Scheme of Examination

Year / Semester	Number, Code or ID and Nomenclature of Paper			Duration of Exam. (in Hrs.)	Teaching Hrs / Week & Credit Points			Distribution of Assessment Marks				Total Marks	
	Number of Paper	Code or ID of Paper	Nomenclature of Paper		Teaching		Credit Points	Continuous Assessment (20%)		Semester Assessment (80%)			
					Th.	Pr.		Max. Marks	Min. Pass Marks	Max. Marks	Min. Pass Marks	Max. Marks	Min. Pass Marks
2nd Year III Semester	Paper-3.1	CHEM-231	Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-3.2	CHEM-232	Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-3.3	CHEM-233	Chemistry Practical	6	-	4	2	--	--	50	25	50	25
	Paper-3.4	MATH-231	Mathematics-I	3	3	-	3	15	06	60	24	75	30
	Paper-3.5	MATH-232	Mathematics-II	3	3	-	3	15	06	60	24	75	30
	Paper-3.6	MATH-233	Mathematics Practical	6	-	4	2	--	--	50	25	50	25
	Paper-3.7	PHY-231	Physics-I	3	3	-	3	15	06	60	24	75	30
	Paper-3.8	PHY-232	Physics-II	3	3	-	3	15	06	60	24	75	30
	Paper-3.9	PHY-233	Physics Practical	6	-	4	2	--	--	50	25	50	25
Total (III Semester)				36	30	24	90	36	510	219	600	195	
2nd Year IV Semester	Paper-4.1	CHEM-241	Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-4.2	CHEM-242	Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-4.3	CHEM-243	Chemistry Practical	6	-	4	2	--	--	50	25	50	25
	Paper-4.4	MATH-241	Mathematics-I	3	3	-	3	15	06	60	24	75	30
	Paper-4.5	MATH-242	Mathematics-II	3	3	-	3	15	06	60	24	75	30
	Paper-4.6	MATH-243	Mathematics Practical	6	-	4	2	--	--	50	25	50	25
	Paper-4.7	PHY-241	Physics-I	3	3	-	3	15	06	60	24	75	30
	Paper-4.8	PHY-242	Physics-II	3	3	-	3	15	06	60	24	75	30
	Paper-4.9	PHY-243	Physics Practical	6	-	4	2	--	--	50	25	50	25
Total (IV Semester)				36	30	24	90	36	510	219	600	195	

Bachelor of Science (B.Sc.) (Mathematics Group)

Subject Combination: Chemistry, Mathematics, Physics (CMP)

B.Sc. 3rd Year : Fifth and Sixth Semesters Semester Scheme of Examination

Year / Semester	Number, Code or ID and Nomenclature of Paper			Duration of Exam. (in Hrs.)	Teaching Hrs / Week & Credit Points			Distribution of Assessment Marks				Total Marks	
	Number of Paper	Code or ID of Paper	Nomenclature of Paper		Teaching Th.	Pr.	Credit Points	Continuous Assessment (20%)		Semester Assessment (80%)			
								Max. Marks	Min. Pass Marks	Max. Marks	Min. Pass Marks	Max. Marks	Min. Pass Marks
3rd Year V Semester	Paper-5.1	CHEM-351	Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-5.2	CHEM-352	Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-5.3	CHEM-353	Chemistry Practical	6	-	4	2	--	--	50	25	50	25
	Paper-5.4	MATH-351	Mathematics-I	3	3	-	3	15	06	60	24	75	30
	Paper-5.5	MATH-352	Mathematics-II	3	3	-	3	15	06	60	24	75	30
	Paper-5.6	MATH-353	Mathematics Practical	6	-	4	2	--	--	50	25	50	25
	Paper-5.7	PHY-351	Physics-I	3	3	-	3	15	06	60	24	75	30
	Paper-5.8	PHY-352	Physics-II	3	3	-	3	15	06	60	24	75	30
	Paper-5.9	PHY-353	Physics Practical	6	-	4	2	--	--	50	25	50	25
Total (V Semester)				36	30	24	90	36	510	219	600	195	
3rd Year VI Semester	Paper-6.1	CHEM-361	Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-6.2	CHEM-362	Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-6.3	CHEM-363	Chemistry Practical	6	-	4	2	--	--	50	25	50	25
	Paper-6.4	MATH-361	Mathematics-I	3	3	-	3	15	06	60	24	75	30
	Paper-6.5	MATH-362	Mathematics-II	3	3	-	3	15	06	60	24	75	30
	Paper-6.6	MATH-363	Mathematics Practical	6	-	4	2	--	--	50	25	50	25
	Paper-6.7	PHY-361	Physics-I	3	3	-	3	15	06	60	24	75	30
	Paper-6.8	PHY-362	Physics-II	3	3	-	3	15	06	60	24	75	30
	Paper-6.9	PHY-363	Physics Practical	6	-	4	2	--	--	50	25	50	25
Total (VI Semester)				36	30	24	90	36	510	219	600	195	

Bachelor of Science (B.Sc.)
(Biology Group)
Subject Combination: Chemistry, Botany, Zoology (CBZ)
B.Sc. 1st Year: First and Second Semesters
Semester Scheme of Examination

Year / Semester	Number, Code or ID and Nomenclature of Paper			Duration of Exam. (in Hrs.)	Teaching Hrs / Week & Credit Points			Distribution of Assessment Marks				Total Marks	
	Number of Paper	Code or ID of Paper	Nomenclature of Paper		Teaching Th.	Pr.	Credit Points	Continuous Assessment (20%)		Semester Assessment (80%)			
								Max. Marks	Min. Pass Marks	Max. Marks	Min. Pass Marks		
1st Year I Semester	Paper-1.1	HIND-111	General Hindi	2	2	-	2	-	-	50	20	50	20
	Paper-1.2	ECA-112	Elementary Computer Applications	2	2	-	2	-	-	50	20	50	20
	Paper-1.3	CHEM-111	Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-1.4	CHEM-112	Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-1.5	CHEM-113	Chemistry Practical	6	-	4	2	--	--	50	25	50	25
	Paper-1.6	BOT-111	Botany-I	3	3	-	3	15	06	60	24	75	30
	Paper-1.7	BOT-112	Botany-II	3	3	-	3	15	06	60	24	75	30
	Paper-1.8	BOT-113	Botany Practical	6	-	4	2	--	--	50	25	50	25
	Paper-1.9	ZOO-111	Zoology-I	3	3	-	3	15	06	60	24	75	30
	Paper-1.10	ZOO-112	Zoology-II	3	3	-	3	15	06	60	24	75	30
	Paper-1.11	ZOO-113	Zoology Practical	6	-	4	2	--	--	50	25	50	25
Total (I Semester)					30+4		24+4	90	36	510+100	219+40	600+100	195+40
1st Year II Semester	Paper-2.1	ENG-121	General English	2	2	-	2	-	-	50	20	50	20
	Paper-2.2	ENV-122	Environmental Studies	2	2	-	2	-	-	50	20	50	20
	Paper-2.3	CHEM-121	Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-2.4	CHEM-122	Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-2.5	CHEM-123	Chemistry Practical	6	-	4	2	--	--	50	25	50	25
	Paper-2.6	BOT-121	Botany-I	3	3	-	3	15	06	60	24	75	30
	Paper-2.7	BOT-122	Botany-II	3	3	-	3	15	06	60	24	75	30
	Paper-2.8	BOT-123	Botany Practical	6	-	4	2	--	--	50	25	50	25
	Paper-2.9	ZOO-121	Zoology-I	3	3	-	3	15	06	60	24	75	30
	Paper-2.10	ZOO-122	Zoology-II	3	3	-	3	15	06	60	24	75	30
	Paper-2.11	ZOO-123	Zoology Practical	6	-	4	2	--	--	50	25	50	25
Total (II Semester)					30+4		24+4	90	36	510+100	219+40	600+100	195+40

Bachelor of Science (B.Sc.) (Biology Group)

Subject Combination: Chemistry, Botany, Zoology (CBZ)

B.Sc. 2nd Year: Third and Fourth Semesters Semester Scheme of Examination

Year / Semester	Number, Code or ID and Nomenclature of Paper			Duration of Exam. (in Hrs.)	Teaching Hrs / Week & Credit Points			Distribution of Assessment Marks				Total Marks	
	Number of Paper	Code or ID of Paper	Nomenclature of Paper		Teaching Th.	Pr.	Credit Points	Continuous Assessment (20%)		Semester Assessment (80%)			
								Max. Marks	Min. Pass Marks	Max. Marks	Min. Pass Marks	Max. Marks	Min. Pass Marks
2nd Year III Semester	Paper-3.1	CHEM-231	Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-3.2	CHEM-232	Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-3.3	CHEM-233	Chemistry Practical	6	-	4	2	--	--	50	25	50	25
	Paper-3.4	BOT-231	Botany-I	3	3	-	3	15	06	60	24	75	30
	Paper-3.5	BOT-232	Botany-II	3	3	-	3	15	06	60	24	75	30
	Paper-3.6	BOT-233	Botany Practical	6	-	4	2	--	--	50	25	50	25
	Paper-3.7	ZOO-231	Zoology-I	3	3	-	3	15	06	60	24	75	30
	Paper-3.8	ZOO-232	Zoology-II	3	3	-	3	15	06	60	24	75	30
	Paper-3.9	ZOO-233	Zoology Practical	6	-	4	2	--	--	50	25	50	25
Total (III Semester)				36	30	24	90	36	510	219	600	195	
2nd Year IV Semester	Paper-4.1	CHEM-241	Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-4.2	CHEM-242	Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-4.3	CHEM-243	Chemistry Practical	6	-	4	2	--	--	50	25	50	25
	Paper-4.4	BOT-241	Botany-I	3	3	-	3	15	06	60	24	75	30
	Paper-4.5	BOT-242	Botany-II	3	3	-	3	15	06	60	24	75	30
	Paper-4.6	BOT-243	Botany Practical	6	-	4	2	--	--	50	25	50	25
	Paper-4.7	ZOO-241	Zoology-I	3	3	-	3	15	06	60	24	75	30
	Paper-4.8	ZOO-242	Zoology-II	3	3	-	3	15	06	60	24	75	30
	Paper-4.9	ZOO-243	Zoology Practical	6	-	4	2	--	--	50	25	50	25
Total (IV Semester)				36	30	24	90	36	510	219	600	195	

Bachelor of Science (B.Sc.)
(Biology Group)
Subject Combination: Chemistry, Botany, Zoology (CBZ)

B.Sc. 3rd Year: Fifth and Sixth Semesters
Semester Scheme of Examination

Year / Semester	Number, Code or ID and Nomenclature of Paper			Duration of Exam. (in Hrs.)	Teaching Hrs / Week & Credit Points			Distribution of Assessment Marks				Total Marks	
	Number of Paper	Code or ID of Paper	Nomenclature of Paper		Teaching Th.	Pr.	Credit Points	Continuous Assessment (20%)		Semester Assessment (80%)			
								Max. Marks	Min. Pass Marks	Max. Marks	Min. Pass Marks	Max. Marks	Min. Pass Marks
3rd Year V Semester	Paper-5.1	CHEM-351	Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-5.2	CHEM-352	Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-5.3	CHEM-363	Chemistry Practical	6	-	4	2	--	--	50	25	50	25
	Paper-5.4	BOT-351	Botany-I	3	3	-	3	15	06	60	24	75	30
	Paper-5.5	BOT-352	Botany-II	3	3	-	3	15	06	60	24	75	30
	Paper-5.6	BOT-353	Botany Practical	6	-	4	2	--	--	50	25	50	25
	Paper-5.7	ZOO-351	Zoology-I	3	3	-	3	15	06	60	24	75	30
	Paper-5.8	ZOO-352	Zoology-II	3	3	-	3	15	06	60	24	75	30
	Paper-5.9	ZOO-353	Zoology Practical	6	-	4	2	--	--	50	25	50	25
Total (V Semester)				36	30	24	90	36	510	219	600	195	
3rd Year VI Semester	Paper-6.1	CHEM-361	Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-6.2	CHEM-362	Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-6.3	CHEM-363	Chemistry Practical	6	-	4	2	--	--	50	25	50	25
	Paper-6.4	BOT-361	Botany-I	3	3	-	3	15	06	60	24	75	30
	Paper-6.5	BOT-362	Botany-II	3	3	-	3	15	06	60	24	75	30
	Paper-6.6	BOT-363	Botany Practical	6	-	4	2	--	--	50	25	50	25
	Paper-6.7	ZOO-361	Zoology-I	3	3	-	3	15	06	60	24	75	30
	Paper-6.8	ZOO-362	Zoology-II	3	3	-	3	15	06	60	24	75	30
	Paper-6.9	ZOO-363	Zoology Practical	6	-	4	2	--	--	50	25	50	25
Total (VI Semester)				36	30	24	90	36	510	219	600	195	

Bachelor of Science (B.Sc.) CHEMISTRY

(Common for Mathematics / Biology Group: All Semesters)

Subject Combination: Chemistry, Physics, Mathematics (CPM) / Chemistry, Botany, Zoology (CBZ)

Semester Scheme of Examination

Year / Semester	Number, Code or ID and Nomenclature of Paper			Duration of Exam. (in Hrs.)	Teaching Hrs / Week & Credit Points			Distribution of Assessment Marks				Total Marks	
	Number of Paper	Code or ID of Paper	Nomenclature of Paper		Teaching Th.	Pr.	Credit Points	Continuous Assessment (20%)		Semester Assessment (80%)			
								Max. Marks	Min. Pass Marks	Max. Marks	Min. Pass Marks	Max. Marks	Min. Pass Marks
1st Year I Semester	Paper-1.3	CHEM-111	Chemistry-I : Inorganic Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-1.4	CHEM-112	Chemistry-II : Organic Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-1.5	CHEM-113	Practical : Chemistry Practical-I	6	-	4	2	--	--	50	25	50	25
				12	6	4	8	30	12	170	73	200	85
1st Year II Semester	Paper-2.3	CHEM-121	Chemistry-I : Physical Chemistry-I	3	3	-	3	15	06	60	24	75	30
	Paper-2.4	CHEM-122	Chemistry-II : Inorganic Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-2.5	CHEM-123	Practical : Chemistry Practical-II	6	-	4	2	--	--	50	25	50	25
				12	6	4	8	30	12	170	73	200	85
2nd Year III Semester	Paper-3.1	CHEM-231	Chemistry-I : Organic Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-3.2	CHEM-232	Chemistry-II : Physical Chemistry-II	3	3	-	3	15	06	60	24	75	30
	Paper-3.3	CHEM-233	Practical : Chemistry Practical-III	6	-	4	2	--	--	50	25	50	25
				12	6	4	8	30	12	170	73	200	85
2nd Year IV Semester	Paper-4.1	CHEM-241	Chemistry-I : Inorganic Chemistry-III	3	3	-	3	15	06	60	24	75	30
	Paper-4.2	CHEM-242	Chemistry-II : Organic Chemistry-III	3	3	-	3	15	06	60	24	75	30
	Paper-4.3	CHEM-243	Practical : Chemistry Practical-IV	6	-	4	2	--	--	50	25	50	25
				12	6	4	8	30	12	170	73	200	85
3rd Year V Semester	Paper-5.1	CHEM-351	Chemistry-I : Physical Chemistry-III	3	3	-	3	15	06	60	24	75	30
	Paper-5.2	CHEM-352	Chemistry-II : Inorganic Chemistry-IV	3	3	-	3	15	06	60	24	75	30
	Paper-5.3	CHEM-353	Practical : Chemistry Practical-V	6	-	4	2	--	--	50	25	50	25
				12	6	4	8	30	12	170	73	200	85
3rd Year VI Semester	Paper-6.1	CHEM-361	Chemistry-I : Organic Chemistry-IV	3	3	-	3	15	06	60	24	75	30
	Paper-6.2	CHEM-362	Chemistry-II : Physical Chemistry-IV	3	3	-	3	15	06	60	24	75	30
	Paper-6.3	CHEM-363	Practical : Chemistry Practical-VI	6	-	4	2	--	--	50	25	50	25
				12	6	4	8	30	12	170	73	200	85

Rules & Regulations

Objectives of the Course:

Bachelor of Science (B.Sc.) programme of University is a pioneering model in science. The course shall provide the thorough knowledge of all the branches of the chemistry. The course also emphasizes on the communication & presentation skills of the students. After completing the course, the students shall be eligible to take admission for higher studies in different branches of the chemical sciences and able to do research in the different areas chemical sciences or allied fields and shall be placed in different organizations / institutions where skilled chemical science professionals are required.

Duration of the Course:

The course Bachelor of Science (B.Sc.) Pass Course shall consist of three academic years divided in to the six semesters. B.Sc. (Pass Course) degree shall be awarded to the candidates after successful completion of the six semester programme of study.

Eligibility for Admission:

A candidate who has passed 10+2 or equivalent examination with Physics, Chemistry and Mathematics or Physics, Chemistry and Biology from any recognized board shall be permitted to take admission in B.Sc. First Semester to award B.Sc. (Pass Course) degree in Maths. or Bio. group from this University.

Structure of the Programme:

The B.Sc. (Pass Course) programme consists of core and applied courses of theory as well as practical papers which are compulsory for all students.

Course Number, Course Code or ID and Nomenclature:

Number of the Paper has been designated in the Arabic number as Paper-1.1, Paper-1.2, and Paper-1.3 and so on. In the Paper-1.2, 1 represents the semester number and 2 represent the paper number. To give a code to a particular course, following sequence has been adopted:

“Abbreviation of the programme in upper case + nth number of year of study + nth number of semester of the programme + course number in Arabic number”

According to the above sequence, code of paper-IV of the first semester of undergraduate Chemistry shall be as “CHEM-114”. Nomenclature of the particular course has been given according to the nature or type of contents included in the Unit-I to Unit-V of course of study.

Maximum Marks and Credit Points:

Maximum marks of a theory and practical paper shall be decided on the basis of their contact hours / per week. One teaching hour per week shall equal to one credit and carry 25 maximum marks. Therefore, 3 teaching hours or 3 credit points per week shall carry 75 maximum marks for each theory paper/course. While two contact hours per week for a

laboratory or practical work shall be equal to one credit point per week. Therefore, 4 contact hours / week shall equal to 2 credit points per week and shall carry 50 maximum marks.

Attendance:

Every teaching faculty, handling a course, shall be responsible for the maintenance of Attendance Register for candidates who have registered for the course. The teacher of the course must intimate the Head of the Department at least seven calendar days before the last instruction day in the semester about the attendance particulars of all students. Each student should earn 75% attendance in the courses of the particular semester failing which he or she will not be permitted to sit in the End-Semester Examinations. However, it shall be open to the authorities to grant exemption to a candidate who has failed to obtain the prescribed 75% attendance for valid reasons and such exemptions should not under any circumstance be granted for attendance below 65%.

Teaching Methodologies:

The classroom teaching would be through conventional lectures or use of OHP or power point presentations (PPT). The lecture would be such that the student should participate actively in the discussion. Student seminars would be conducted and scientific discussions would be arranged to improve their communicative skill. In the laboratory, instruction would be given for the experiments followed by demonstration and finally the students have to do the experiments individually. For the students of slow learners, special attention would be given.

Assessment Pattern / Scheme of Examination:

The assessment of the student shall be divided into two parts in which first part is continuous assessment or internal assessment or mid-term assessment (20% weightage of the maximum marks) and second part is semester assessment or external assessment or end-term assessment (80% weightage of the maximum marks).

(i) Continuous Assessment or Internal or Mid Term Assessment:

- (a) The continuous or internal or mid-term assessment (20% weightage of the maximum marks) for each theory paper shall be taken by the faculty members in the Department during each semester. There will be two internal assessment tests (*i.e.* first internal assessment test or first mid-term test and second internal assessment test or second mid-term test) each of 10% weightage of maximum marks of each theory paper. Each internal assessment test shall be of one hour duration for theory paper and shall be taken according to academic calendar which will be notified by the Department / University.
- (b) For practical papers, there will be only one external or semester or end term assessment (100% weightage of maximum marks) and there will be no continuous or internal or midterm assessment.

- (c) A student who remains absent (defaulter) or fails or wants to improve the marks in the internal assessment may be permitted to appear in the desired paper(s) (only one time) in the same semester with the permission of the concern Head of the Department. A defaulter / improvement fee of Rupees 250/- per paper shall be taken from such candidates. Duly forwarded application of such candidates by the teacher concerned shall be submitted to Head of the Department who may permit the candidate to appear in the internal assessment after production of satisfactory evidence about the reason of his/her absence in the test(s) and deposition of the defaulter / improvement fee. A record of such candidates shall be kept in the Department.
- (d) Regular attendance of the student shall be considered in the internal assessment. Some marks for regularity shall be given to the student(s) who is/are taken classes regularly from the 2% weightage of the maximum marks. The 2% weightage of the maximum marks of regularity shall be taken from the weightage given for second internal assessment (10% weightage of maximum marks). After excluding the 2% weightage of regularity, the second internal assessment shall be of 8% weightage of maximum marks. If the attendance / regularity factor is similar for all the students, then it may be merged with the weightage of second internal assessment test (class test, home assignment, quiz, seminar, *etc.*) and then second internal assessment test shall be of 10% weightage of maximum marks.
- (e) Paper wise consolidated marks for each theory paper and dissertation / seminar (*i.e.* total marks obtained during various modes of internal assessment) obtained by the students (out of the 20% weightage of the maximum marks of the each paper) shall be forwarded by the Head of the Department (in two copies) to the Controller of Examinations of the University within a week from the date of last internal assessment test for incorporation in the tabulation register.
- (f) The consolidated marks obtained by the students be also made known to them before being communicated by the concerned Head of the Department to the University for final incorporation in the tabulation register. If any discrepancies are discovered or pointed out by the students, the same shall be looked into by the concerned faculty member and corrections made wherever necessary. The decision of the Head of the Department before the communication of marks to the University shall be final. No corrections shall be made in the internal assessment marks after the declaration of the result by the University.
- (g) Consolidated marks of internal assessment obtained out of the 20% weightage of maximum marks of each theory paper which will be communicated to the University shall be in whole number and not in fraction. Marks awarded for the various internal assessments in each paper shall be added up and then round off to the next whole number to avoid any fraction.

- (h) All test copies and other material related to the internal assessment shall also be sent to the Controller of Examinations of the University to keep in record as per the University guidelines.
- (i) The concerned Head of the Department shall be responsible for proper conduct of internal assessment tests and for communication of the consolidated marks to the University within the prescribed time.
- (j) The Head of the Department shall keep a record of the marks and also notify the same to the candidates immediately so that if any candidate is not satisfied with the award in any test or seasonal work, he / she should represent the matter to the higher authority.

(ii) Semester Assessment or External or End Term Assessment:

- (a) The semester or external or end-term assessment (80% weightage of the maximum marks) shall be three hours duration to each theory paper and six hours duration for each practical paper and shall be taken by the University at the end of each semester.
- (b) The syllabus for each theory paper is divided into five independent units.

Question Paper Pattern:

(A) Continuous or Internal or Mid Term Assessment:

(i) First Continuous or Internal or Mid Term Assessment:

Department of Pure & Applied Chemistry
University of Kota
Kota (Rajasthan)-324 005

First Internal Assessment Test 20... - 20....

Class	:	Max. Marks	: 7.5
Semester	:	No. of Students	:
Subject	:	Duration of Exam	:
Paper	:	Name of Teacher	:

Q. No. 1.
or

2.5 Marks

Q. No. 2.
or

2.5 Marks

Q. No. 3.
or

2.5 Marks

(ii) Second Continuous or Internal or Mid Term Assessment:

(a) Attendance:

Marks shall be given by the faculty member in each paper according to its weightage.

1.5 (2% weightage of Maximum Marks)

Note:

If the attendance / regularity factor is similar for all the students, then it may be merged with the weightage of second internal assessment test (class test / home assignment / quiz, seminar, etc.).

(b) Class Test:

6 (8 % weightage of Maximum Marks)

Format

Department of Pure & Applied Chemistry
University of Kota
Kota (Rajasthan)-324 005
Second Internal Assessment Test 20... - 20....

Class	:	Max. Marks	: 6
Semester	:	No. of Students	:
Subject	:	Duration of Exam	:
Paper	:	Name of Teacher	:

Q. No. 1.
2 Marks

Q. No. 2.
or

.....
2 Marks

Q. No. 3.
or

.....
2 Marks

or

(b) Assignment:

(May be divided in parts or questions or may not be. It will be depending on the nature of assignment).

6 (8 % weightage of Maximum Marks)

or

(b) Quiz:

(May be divided in parts or questions or may not be. It will be depending on the nature of quiz).

6 (8 % weightage of Maximum Marks)

or

(b) Any other tool may be adopted for internal Assessment

6 (8 % weightage of Maximum Marks)

(B) Semester or External or End Term Assessment:

Duration of Examination: 3 Hours

Max. Marks: 60

There will be ten long answer type questions covering all units but not more than two questions from each unit, descriptive type and answer in about 400 words. Students have to attempt 5 questions taking one from each unit. Paper setter shall be advised to frame the two questions from each unit covering all five units. All the questions will carry equal marks.

	Unit-I	
Q. No. 1		12 Marks
	or	
Q. No. 2		12 Marks
	Unit-II	
Q. No. 3		12 Marks
	or	
Q. No. 4		12 Marks
	Unit-III	
Q. No. 5		12 Marks
	or	
Q. No. 6		12 Marks
	Unit-IV	
Q. No. 7		12 Marks
	or	
Q. No. 8		12 Marks
	Unit-V	
Q. No. 9		12 Marks
	or	
Q. No. 10		12 Marks

Practical Examinations:

Continuous or Internal or Mid Term Assessment: *Not applicable in practical.*

External or Semester or End Term Assessment:

Duration of Exam: 6 Hours

Maximum Marks: 50

Distribution of Maximum Marks:

S. No.	Name of Exercise	Marks
1.	Exercise No. 1	15
2.	Exercise No. 2	15
3.	Viva-voce	10
4.	Practical Record	10
Total Marks		50

Minimum Pass Marks and Rules regarding Determination of Results:

Each semester shall be regarded as a unit for working out the result of the candidates. The result of the each semester examination shall be worked out separately (even if he/she has appeared at the paper of the lower semester along with the papers of higher semester) in accordance with the following conditions:

- (a) The candidate shall be declared as pass in a semester examination, if he/she secures at least 40% marks in each theory paper separately in external & internal examination and 50% marks in each practical paper / project / dissertation with 40% aggregate marks in that semester.
- (b) A candidate declared as fail/absent in one or more papers at any odd semester examination shall be permitted to take admission in the next higher semester (even semester) of the same academic session.
- (c) A candidate may be promoted in the next academic session (odd semester) if he/she has cleared collectively at least 50% of the papers of both semesters of previous academic session. The candidate who does not fulfill the above condition will remain as an ex-student and will reappear in the due papers along with next odd/even semester exams.
- (d) If any student who is provisionally admitted in higher odd semester but could not secure prescribed minimum marks in previous semesters will be treated as ex-student and his/her admission fee will be carry forwarded to the next odd semester of forthcoming academic session.
- (e) If a candidate, who is declared as pass, wishes to improve his/her performance in the theory papers of previous semester, he/she may re-appear only one time in these papers in next odd/even semester examinations.
- (f) Candidate shall not be permitted to re-appear or improve the marks obtained in the external examination of practical / dissertation in any condition.
- (g) If the number of papers prescribed in a semester examination is an odd number, it shall be increased by one for the purpose of reckoning 50% of the papers for considering the student pass/fail.
- (h) A candidate may be given only two additional chances for passing the semester thus maximum tenure for completing for three years under-graduate programme up to five years and so on.
- (i) The marks secured in the Gen Hindi, Gen English, Elementary Computer applications and Environment studies shall not be counted in awarding the division to a candidate. The candidate shall have to clear the compulsory subjects in the additional three chances and non-appearance or absence in the examination of compulsory subjects shall be counted as chance and shall be declared fail in that examination.
- (j) The grace marks scheme shall be applicable as per University norms.

Classification of Successful Candidates:

The classification of successful candidates after last semester examination shall be as under:

Description of Marks Obtained	Division / Result
• 75% and above marks in a paper.	Distinction in that paper
• A candidate who has secured aggregate 60% and above marks	First Division
• A candidate who has secured aggregate 50% and above but less than 60% marks	Second Division
• A candidate who has secured aggregate 40% and above but less than 50% marks	Pass

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Syllabus

B.Sc. Chemistry First Semester Examination

Paper-1.3: CHEM-111: Inorganic Chemistry-I

Contact Hours / Week	: 3 Hours / Week	Maximum Marks	: 75 Marks
Duration of Examination	: 3 Hours	Continuous Assessment	: 15 Marks
		Semester Assessment	: 60 Marks

The syllabus is divided into five independent units and There will be ten long answer type questions covering all units but not more than two questions from each unit, descriptive type and answer in about 400 words. Students have to attempt 5 questions taking one from each unit. Paper setter shall be advised to frame the two questions from each unit covering all five units. All the questions will carry equal marks.

Note: Contents of each unit may be completed into 8-10 lectures or contact hours which also include revisions, seminars, internal assessments, etc.

Unit-I Development of Chemical Sciences in India:

Alchemical ideas in Vedas, ideas about ayurvedic, transitional, tantric and iatro-chemical periods, constitution and properties of matter, metals and metallurgy, discovery of medicines, food technology, fermentation technology, textile technology, building materials, Indigenous origin of Indian alchemy, some noted Indian alchemists and their works.

Unit-II Atomic Structure:

Idea of De Broglie matter waves, Heisenberg's uncertainty principle, atomic orbitals, Schrodinger wave equation, significance of ψ and ψ^2 , quantum numbers, radial and angular wave function and probability distribution curves, shapes of s, p, d, orbitals, Aufbau and Pauli exclusion principles, Hund's multiplicity rule, electronic configurations of the elements, effective nuclear charge.

Unit-III Periodic Properties:

Atomic and ionic radii, ionization energy, electron affinity and electronegativity-definition, methods of determination or evaluation, trends in periodic table and applications in predicting and explaining the chemical behaviour.

Unit-IV Chemical Bonding-I:

Covalent Bond: Valence bond theory and its limitations, directional characteristics of covalent bond, various types of hybridizations and shapes of simple inorganic molecules and ions. Valence shell electron pair repulsion (VSEPR) theory to NH_3 , H_3O^+ , SF_4 , ClF_3 , ICl_2 and H_2O ; MO theory, homonuclear and heteronuclear (CO and NO) diatomic molecules, multicenter bonding in electron deficient molecules, bond

strength and bond energy, percentage ionic character from dipole moment and electronegativity difference.

Unit-V Chemical Bonding-II:

Ionic Solids: Ionic structure, radius ratio effect and coordination number, limitations of radius ratio rule, lattice defects, semi-conductors, lattice energy and Born-Haber cycle, solvation energy and solubility of ionic solids, polarizing power and polarizability of ions, Fajan's rule, Metallic bond: Free electron, valence bond and band theories.

Weak Interactions: Hydrogen bonding, van der Waals forces.

Books suggested:

- *Advanced Inorganic Chemistry, Vol I & II. Satya Praksh, G.D. Tuli, S.K. Basu and R.D. Madan*
- *Principles of Inorganic Chemistry : B. R. Puri and L. R. Sharma*
- *Shriver & Atkins' Inorganic Chemistry*
- *Concise Inorganic Chemistry : J. D. Lee*
- *General Inorganic Chemistry: J. A. Duffly, Longman (2nd Ed.)*
- *Basic Inorganic Chemistry : F. A. Cotton and G. Wilkinson, Wiley Eastern*

Paper-1.4: CHEM-112: Organic Chemistry-I

Contact Hours / Week	: 3 Hours / Week	Maximum Marks	: 75 Marks
Duration of Examination	: 3 Hours	Continuous Assessment	: 15 Marks
		Semester Assessment	: 60 Marks

The syllabus is divided into five independent units and There will be ten long answer type questions covering all units but not more than two questions from each unit, descriptive type and answer in about 400 words. Students have to attempt 5 questions taking one from each unit. Paper setter shall be advised to frame the two questions from each unit covering all five units. All the questions will carry equal marks.

Note: Contents of each unit may be completed into 8-10 lectures or contact hours which also include revisions, seminars, internal assessments, etc.

Unit-I Structure and Bonding:

Hybridization, bond lengths and bond angles, bond energy, localized and delocalized chemical bonds, van der Waals interactions, inclusion compounds, clathrates, charge transfer complexes, resonance, hyperconjugation, inductive and field effects, hydrogen bonding.

Unit-II Mechanism of Organic Reactions:

Curved arrow notation, drawing electron movements with arrows, half headed and double headed arrow, homolytic and heterolytic bond breaking. Types of reagents- electrophiles and nucleophiles. Types of organic reactions, Energy considerations.

Reactive Intermediates:

Carbocations, carbanions, free radicals, carbenes, arynes and nitrenes (with examples). Assigning formal charges on intermediates and other ionic species.

Methods of determination of reaction mechanism (product analysis, intermediates, isotope effects. Kinetic and stereochemical studies).

Unit-III Stereochemistry of Organic Compounds:

Concept of isomerism, types of isomerism. *Optical isomerism:* Elements of symmetry, molecular chirality, enantiomers, stereogenic centre, optical activity, properties of enantiomers, chiral and achiral molecules with two stereogenic centres, diastereomers, threo and erythro diastereomers, meso compounds, resolution of enantiomers, inversion, retention and racemization. Relative and absolute configuration: Sequence rules, D&L and R&S systems of nomenclature. *Geometric isomerism:* Determination of configuration of geometric isomers, E&Z systems of nomenclature, geometric isomerism in oximes and alicyclic compounds. *Conformational isomerism:* Conformational analysis of ethane and n-butane; conformations of cyclohexane, axial and equatorial bonds, conformation of mono substituted cyclohexane derivatives. Newmann projection and sawhorse formulae, Fischer and Flying wedge formulae. Difference between configuration and conformation.

Unit-IV Alkanes:

IUPAC nomenclature of branched and unbranched alkanes, alkyl group, classification of carbon atoms in alkanes, isomerism in alkanes, sources, methods of preparation (with special reference to Wurtz reaction, Kolbe reactions, Corey-House reaction and decarboxylation of carboxylic acids), physical properties and chemical reaction of alkanes. Mechanism of free radical halogenation of alkanes: orientation, reactivity and selectivity.

Cycloalkanes:

Nomenclature, methods of preparation, chemical reactions, Baeyer's strain theory and its limitations, Ring strains in small rings (cyclopropane and cyclobutane), theory of strainless rings. The case of cyclopropane ring: banana bonds.

Unit-V Alkenes:

Nomenclature of alkenes, methods of preparation, mechanisms of dehydration of alcohols and dehydrohalogenation of alkyl halides, regioselectivity in alcohol dehydration, the Saytzeff rule, Hofmann elimination, physical properties and relative stabilities of alkenes. Chemical reactions of alkenes-Mechanisms involved in hydrogenation, electrophilic and free radical additions, Markownikof's rule, hydroboration-oxidation, oxymercuration-reduction, epoxidation, ozonolysis, hydration, hydroxylation and oxidation with KMnO_4 , polymerization of alkenes. Substitution at the allylic and vinylic positions of alkenes. Industrial applications of ethylene and propene.

Cycloalkenes and Dienes:

Methods of formation, conformation and chemical reactions of cycloalkenes. Nomenclature and classification of dienes: isolated, conjugated and cumulated dienes. Structure of allenes and butadiene, methods of preparation, polymerization. Chemical reactions-1,2- and 1,4-additions, Diels-Alder reaction.

Alkynes:

Nomenclature, structure and bonding in alkynes. Methods of preparation. Chemical reactions of alkynes, acidity of alkynes. Mechanisms of electrophilic and nucleophilic addition reactions, hydroboration-oxidation, metal-ammonia reduction, oxidation and polymerization.

Books Suggested :

- *Structure and Mechanism in Organic Chemistry*, C.K. Ingold, Cornell University Press.
- *Organic Chemistry*, Clayden, Nick Greeves and Stuart Warren, Oxford University Press
- *A Guide Book to Mechanism in Organic Chemistry*, Peter Sykes, Longman.
- *Textbook of Organic Chemistry* by P S Kalsi, New Age International
- *Stereochemistry of Carbon Compounds*, Ernest L. Eliel, TataMcGraw Hill.
- *Stereochemistry of Organic Compounds*, D. Nasipuri, New Age International.
- *Stereochemistry of Organic Compounds*, P.S. Kalsi, New Age International.
- *Organic Chemistry*, R. T. Morrison and R. N. Boyd, Prentice-Hall.
- *Organic Chemistry*, Vol. I, II & III. Jag Mohan, R. Chand & Company
- *Organic Chemistry*, (Vol. I, II & III). S. M. Mukherji, S. P. Singh and R. P. Kapoor
- *A Text Book of Organic Chemistry* : B. S. Bahl and Arun Bahl
- *A Text Book of Organic Chemistry* : P. L. Soni & H.M. Chawla
- *A Text Book of Organic Chemistry: (Vol. I & II)* O. P. Agarwal
- *Modern Principles of Organic Chemistry* : M. K. Jain and S. C. Sharma.

Paper-1.5: CHEM-113: Chemistry Practical-I

Contact Hours / Week : 4 Hours / Week Maximum Marks : 50 Marks
Duration of Examination : 6 Hours Semester Assessment : 50 Marks

Distribution of Marks:

S. No.	Name of Exercise	Marks
1.	Exercise No. 1: Inorganic Chemistry Experiment	15
2.	Exercise No. 2: Organic Chemistry Experiment	15
3.	Viva-voce	10
4.	Practical Record	10
Total Marks		50

Inorganic Chemistry:

Semi-micro / macro Analysis:

Anion analysis (3 radicals).

Cation analysis: Separation and identification of ions from groups I, II, III, IV, V and VI (3 radicals). (Total 6 radicals).

Organic Chemistry:

Laboratory Techniques:

Section-A

- Determination of melting point
 - Naphthalene 80-82°C, Benzoic acid 121.5-122 °C, Urea 132.5-133°C, Succinic Acid 184.5-185°C, Cinnamic acid 132.5-133°C, Salicylic acid 157.5-158°C, Acetanilide 113.5-114 °C, *m*-Dinitrobenzene 90°C, *p*-Dichlorobenzene 52°C, Aspirin 135°C
- Determination of boiling points
 - Ethanol 78°C, Cyclohexane 81.4°C, Toluene 110.6°C, Benzene 80°C
- Determination of mixed melting point
 - Urea-Cinnamic acid mixture of various compositions (1:4,1:1,4:1)
- Distillation
 - Simple distillation of ethanol-water using water condenser
 - Distillation of nitrobenzene and aniline using air condenser
 - Steam Distillation:
 - Naphthalene from its suspension in water.
 - Clove oil from Clove
 - Separation of *o*- and *p*-nitrophenols
- Sublimation (Simple and vacuum)
 - Camphor, Naphthalene, phthalic acid and Succinic acid.
- Crystallization
 - Concept of induction of crystallization.
 - Phthalic acid from hot water (using fluted filter paper and stemless funnel).
 - Acetanilide from boiling water.
 - Naphthalene from Ethanol.
 - Benzoic acid from water.
- Decolorisation and crystallization using charcoal
 - Decolorisation of brown sugar (sucrose) with animal charcoal using gravity filtration.
 - Crystallization and decolorisation of impure naphthalene (100g of naphthalene mixed with 0.3g. of Congo Red using 1.0g decolorising carbon) from ethanol.

Section-B

- Qualitative Analysis:
Detection of extra elements (N, S and halogens) and functional groups (phenolic, carboxylic, carbonyl, ester, carbohydrates, amine, amide, nitro and anilide) in simple organic compounds.

Books Suggested :

- *Vogel's Textbook of Quantitative Analysis*, J. Bassett, R.C. Denney, G.H. Jeffery and J. Mendham
- *Macro scale and Micro scale Organic Experiments*, K.L. Williamson, D.C. Heath.
- *Vogel's Textbook of Practical Organic Chemistry*, A.R. Tatchell, John Wiley.
- *Systematic Qualitative Organic Analysis*, H. Middleton, Edward Arnold.
- *Handbook of Organic Analysis: Qualitative and Quantitative*. H. Clark, Edward Arnold.
- *Experiments and Techniques in Organic Chemistry*, D.P. Pasto, Johnson and Miller, Prentice Hall.

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Syllabus

B.Sc. Chemistry **Second Semester Examination**

Paper-2.3: CHEM-121: Physical Chemistry-I

Contact Hours / Week	: 3 Hours / Week	Maximum Marks	: 75 Marks
Duration of Examination	: 3 Hours	Continuous Assessment	: 15 Marks
		Semester Assessment	: 60 Marks

The syllabus is divided into five independent units and There will be ten long answer type questions covering all units but not more than two questions from each unit, descriptive type and answer in about 400 words. Students have to attempt 5 questions taking one from each unit. Paper setter shall be advised to frame the two questions from each unit covering all five units. All the questions will carry equal marks.

Note: Contents of each unit may be completed into 8-10 lectures or contact hours which also include revisions, seminars, internal assessments, etc.

Unit-I Gaseous State:

Postulates of kinetic theory of gases, deviation from ideal behaviour, van der Waals equation of state.

Critical Phenomena: PV isotherms of real gases, continuity of states, the isotherms of van der Waals equation, relationship between critical constants and van der Waals constants, the law of corresponding states, reduced equation of state.

Molecular Velocities: Root mean square, average and most probable velocities. Qualitative discussions of the Maxwell's distribution of molecular velocities, collision number, mean free path and collision diameter. Liquification of gases (based on Joule-Thomson effect).

Unit II Liquid State:

Intermolecular forces, structure of liquids (a qualitative description). Structural differences between solids, liquids and gases.

Liquid Crystals: Difference between liquid crystal, solid and liquid. Classification, structure of nematic and cholestric phases. Thermography and seven segment cell.

Colloidal State:

Definition of colloids, classification of colloids. Solids in liquids (sols): kinetic, optical and electrical properties; stability of colloids. Protective action, Hardy-Schulze law, gold number. Liquids in liquids (emulsions): types of emulsions, preparation, emulsifier. Liquids in solids (gels): classification, preparation and properties, inhibition, general applications of colloids.

Unit III Solid State:

Classification of solids, definition of space lattice, lattice points, crystal lattice and unit cell. Seven crystal systems. Symmetry elements in crystals. Laws of crystallography (i) Law of constancy of interfacial angles (ii) Law of rationality of

indices (iii) Law of symmetry. X-ray diffraction by crystals. Derivation of Bragg's equation Determination of crystal structure of NaCl, KCl and CsCl (Laue's method and powder method).

Unit IV Chemical Kinetics-I:

Chemical kinetics and its scope, rate of a reaction, factors influencing the rate of a reaction: concentration, temperature, pressure, solvent, light, catalyst. Concentration dependence of rates, mathematical characteristics of simple chemical reactions-zero order, first order, second order, pseudo order, half life and mean life. Determination of the order of reaction: differential method, method of integration, method of half life period and isolation method. Radioactive decay as a first order phenomenon. Experimental methods of chemical kinetics: conductometric, potentiometric, optical methods, polarimetry and spectrophotometry.

Unit V Chemical Kinetics-II:

Theories of chemical kinetics: Effect of temperature on rate of reaction, Arrhenius concept of activation energy. Simple collision theory based on hard sphere model, transition state theory (equilibrium hypothesis). Expression for the rate constant based on equilibrium constant and thermodynamic aspects. Catalysis, characteristics of catalysed reactions, classification of catalysis, miscellaneous examples.

Books Suggested :

- *Principles of Physical Chemistry* : B. R. Puri and L. R. Sharma
- *A Text Book of Physical Chemistry* : A. S. Negi and S. C. Anand
- *Physical Chemistry, Pt. I & II* : C. M. Gupta, J. K. Saxena and M. C. Purohit
- *Atkins' Physical Chemistry*, Oxford University Press.

Paper-2.4: CHEM-122: Inorganic Chemistry-II

Contact Hours / Week	: 3 Hours / Week	Maximum Marks	: 75 Marks
Duration of Examination	: 3 Hours	Continuous Assessment	: 15 Marks
		Semester Assessment	: 60 Marks

The syllabus is divided into five independent units and There will be ten long answer type questions covering all units but not more than two questions from each unit, descriptive type and answer in about 400 words. Students have to attempt 5 questions taking one from each unit. Paper setter shall be advised to frame the two questions from each unit covering all five units. All the questions will carry equal marks.

Note: Contents of each unit may be completed into 8-10 lectures or contact hours which also include revisions, seminars, internal assessments, etc.

Unit-I s-Block Elements:

Comparative study, diagonal relationship, salient features of hydrides, solvation and complexation tendencies including their functions in biosystems, an introduction to alkyls and aryls of s-block elements.

Unit-II p-Block Elements-I:

Comparative study (including diagonal relationship) of groups 13-17 elements, compounds like hydrides, oxides, oxyacids and halides of groups 13-16.

Chemistry of Noble Gases: Chemical properties of Noble gases, chemistry of xenon, structure and bonding in xenon compounds.

Unit-III p-Block Elements-II:

Hydrides of Boron: diborane and higher boranes, borazines, borohydrides, fullerenes, carbides, fluorocarbons, silicates (structural principle), tetrasulphur tetranitride, basic properties of halogens, interhalogens and polyhalides.

Silicones and Phosphazenes: Silicones and phosphazenes as examples of inorganic polymers, nature of bonding in triphosphazenes.

Unit-IV Acids and Bases:

Arrhenius, Bronsted-Lowry, the Lux-Flood solvent system and Lewis concept of acids and bases.

Hard and Soft Acids and Bases (HSAB):

Classification of acids and bases as hard and soft. Pearson's HSAB concept, acid-base strength and hardness and softness. Symbiosis, theoretical basis of hardness and softness, electronegativity and hardness and softness.

Unit-V Non-aqueous Solvents:

Physical properties of solvents, type of solvents and their general characteristics, reactions in liquid NH_3 and Liquid SO_2 .

Oxidation and Reduction:

Use of redox potential data-analysis of redox cycle, redox stability in water, Frost, Latimer and Pourbaix diagrams. Principle involved in the extraction of the elements.

Books suggested:

- *Advanced Inorganic Chemistry, Vol I & II. Satya Prakash, G.D. Tuli, S.K. Basu and R.D. Madan*
- *Principles of Inorganic Chemistry : B. R. Puri and L. R. Sharma*
- *Shriver & Atkins' Inorganic Chemistry*
- *Concise Inorganic Chemistry : J. D. Lee*
- *General Inorganic Chemistry: J. A. Duffy, Longman (2nd Ed.)*
- *Basic Inorganic Chemistry : F. A. Cotton and G. Wilkinson, Wiley Eastern*

Paper-2.5: CHEM-123: Chemistry Practical-II

Contact Hours / Week : 4 Hours / Week Maximum Marks : 50 Marks
Duration of Examination : 6 Hours Semester Assessment : 50 Marks

Distribution of Marks:

S. No.	Name of Exercise	Marks
1.	Exercise No. 1: Inorganic Chemistry Experiment	15
2.	Exercise No. 2: Physical Chemistry Experiment	15
3.	Viva-voce	10
4.	Practical Record	10
Total Marks		50

Inorganic Chemistry:

Semi-micro / macro Analysis:

Anion analysis (3 radicals).

Cation analysis: Separation and identification of ions from groups I, II, III, IV, V and VI (3 radicals).

Physical Chemistry:

Chemical Kinetics

- To determine the specific reaction rate of the hydrolysis of methyl acetate / ethyl acetate catalyzed by hydrogen ions at room temperature.
- To study the effect of acid strength on the hydrolysis of an ester.
- To compare the strengths of HCl and H₂SO₄ by studying the kinetics of hydrolysis of ethyl-acetate.
- To study kinetically the reaction of decomposition of iodide by H₂O₂

Distribution Law

- To study the distribution of iodine between water and CCl₄
- To study the distribution of benzoic acid between benzene and water.

Colloids

- To prepare arsenious sulphide sol and compare the precipitating power of mono-, bi- and trivalent anions.

Viscosity, Surface Tension

- To determine the percentage composition of a given, mixture (non-interacting systems) by viscosity method.
- To determine the viscosity of amyl alcohol in water at different concentrations and calculate the viscosity of these solutions.
- To determine the percentage composition of a given binary mixture by surface tension method (acetone & ethyl methyl ketone).

Books Suggested :

- *Vogel's Textbook of Practical Organic Chemistry*, A.R. Tatchell, John Wiley.
- *Systematic Qualitative Organic Analysis*, H. Middleton, Adward Arnold.
- *Handbook of Organic Analysis: Qualitative and Quantitative*. H. Clark, Adward Arnold.
- *Experiments and Techniques in Organic Chemistry*, D.P. Pasto, Johnson and Miller, Prentice Hall.
- *Practical Chemistry: Giri Bajpai and Pandey*, S. Chand & Co. Ltd., New Delhi.
- *Practical Physical Chemistry*, A.M. James and F.E. Prichard, Longman.
- *Findley's Practical Physical Chemistry*, B.P. Levitt, Longman.
- *Experimental Physical Chemistry*, R.C. Das and B. Behera, Tata McGraw Hill.

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